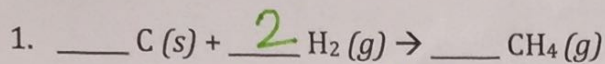


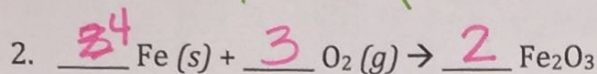
Key

Balancing Chemical Equations Worksheet

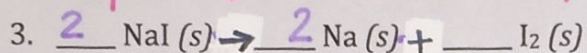
Balance the following reactions by filling in the coefficients.



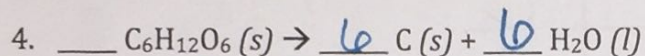
C	H	C	H
1	2	1	4



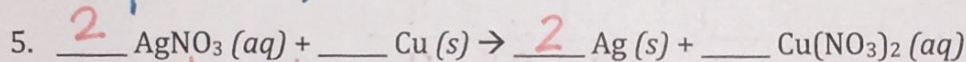
Fe	O	Fe	O
2	6	4	6



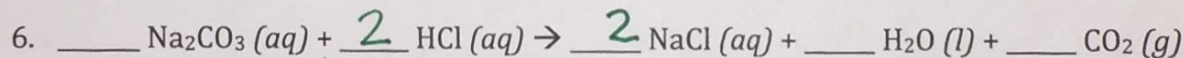
Na	I	Na	I
2	2	2	2



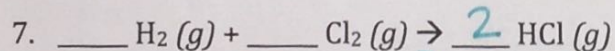
C	H	O	C	H	O
6	12	6	6	12	6



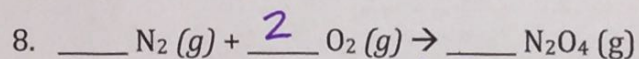
Ag	N	O	Cu	Ag	N	O	Cu
2	2	6	1	2	2	6	1



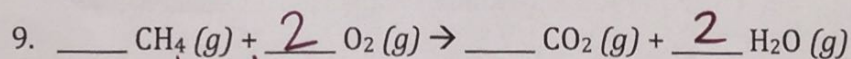
Na	C	O	H	Cl	Na	C	O	H	Cl
2	1	3	2	2	2	1	3	2	2



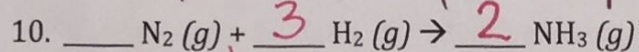
H	Cl	H	Cl
2	2	2	2



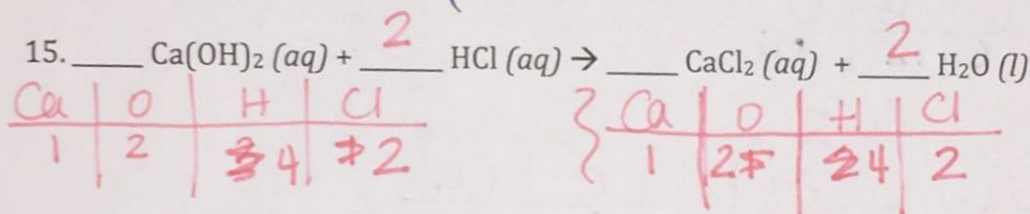
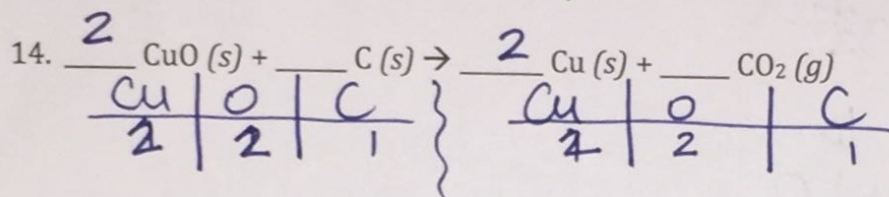
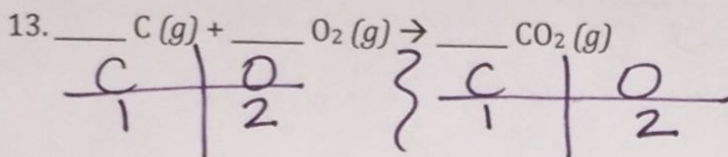
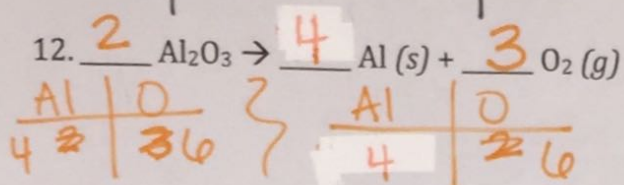
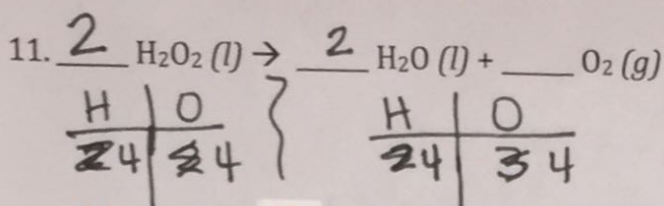
N	O	N	O
2	4	2	4



C	H	O	C	H	O
1	4	4	1	4	4

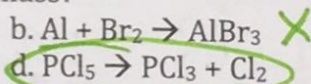
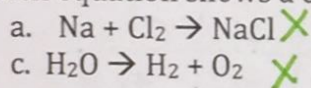


N	H	N	H
2	6	2	6



Regents Practice

16. Which equation shows a conservation of mass?



17. All chemical reactions have a conservation of

a. Mass, only
 c. charge and energy, only

b. mass and charge, only
 d. mass, charge, and energy

18. Given the unbalanced equation: $\text{Fe}_2\text{O}_3 + 3 \text{CO} \rightarrow 2 \text{Fe} + 3 \text{CO}_2$

When the equation is correctly balanced using the *smallest* whole-number coefficients, what is the coefficient of CO ?

- a. 1 b. 2 c. 3 d. 4

19. Consider the following unbalanced equation: $2 \text{Ag} + 1 \text{H}_2\text{S} \rightarrow 1 \text{Ag}_2\text{S} + 1 \text{H}_2$

What is the *sum* of the coefficients when the equation is balanced using the smallest whole-number coefficients?

- a. 5 b. 8 c. 10 d. 4

$2+1+1+1$

20. Given the unbalanced equation: $\text{Mg}(\text{ClO}_3)_2(s) \rightarrow \text{MgCl}_2(s) + 3 \text{O}_2(g)$

What is the coefficient of O_2 when the equation is balanced using the smallest whole-number coefficients?

- a. 1 b. 2 c. 3 d. 4