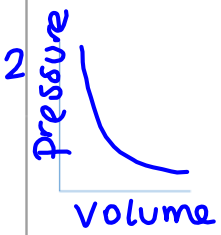
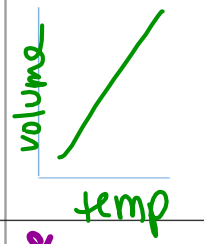
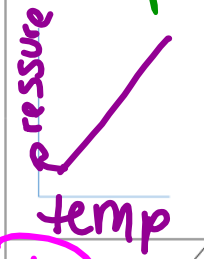
EXPLORING THE GAS LAWS Read pg. 333 -340 and summarize the information in the table below.

Gas Law	Variables	Definition	Equation	Graph	Relationship Circle One
Boyle's Law (pg. 333-335)	Constant Variable: temp. Changing Variables: -volume, - pressure	@ constant temp, as P ↑, V ↓	$P_1V_1 = P_2V_2$		(Inverse) Direct
Charles' Law (pg. 335-337)	Constant Variable: -pressure Changing Variables: -temp. -volume	@ constant pressure, as T ↑, V ↑	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$		Inverse (Direct)
Gay-Lussac's Law (pg. 338)	Constant Variable: -volume Changing Variables: -temp -pressure	@ constant volume, ↑ T, ↑ P	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$		Inverse (Direct)
Combined Gas Laws (pg. 339-340)	Constant Variable: any Changing Variables: any	-combines all 3 gas laws. -the other laws can be obtained by holding one quantity (P, T, or V) constant	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$ table T		