

# IONIC COMPOUNDS - FORMULAS AND NOMENCLATURE WORKSHEET

NAME: \_\_\_\_\_

Write the correct name for:

- |                                     |                           |                                      |                     |
|-------------------------------------|---------------------------|--------------------------------------|---------------------|
| 1) MgS                              | magnesium sulfide         | 16) AlP                              | aluminum phosphide  |
| 2) CaCO <sub>3</sub>                | calcium carbonate         | 17) NaI                              | sodium iodide       |
| 3) Hg <sub>2</sub> F <sub>2</sub>   | dimercury (I) fluoride    | 18) SnCl <sub>2</sub>                | tin (II) chloride   |
| 4) Al <sub>2</sub> O <sub>3</sub>   | aluminum oxide            | 19) HgO                              | mercury (II) oxide  |
| 5) Cu <sub>2</sub> S                | copper (I) sulfide        | 20) Ba <sub>3</sub> N <sub>2</sub>   | barium nitride      |
| 6) SrF <sub>2</sub>                 | strontium fluoride        | 21) AlPO <sub>4</sub>                | aluminum phosphate  |
| 7) Li <sub>2</sub> S                | lithium sulfide           | 22) KNO <sub>2</sub>                 | potassium nitrite   |
| 8) RaCl <sub>2</sub>                | radium chloride           | 23) CaO                              | calcium oxide       |
| 9) NaHCO <sub>3</sub>               | sodium hydrogen carbonate | 24) KBr                              | potassium bromide   |
| 10) Sn <sub>3</sub> P <sub>4</sub>  | tin (IV) phosphide        | 25) Mg(OH) <sub>2</sub>              | magnesium hydroxide |
| 11) CuS                             | copper (II) sulfide       | 26) Na <sub>2</sub> CrO <sub>4</sub> | sodium chromate     |
| 12) PbBr <sub>4</sub>               | lead (IV) bromide         | 27) Ba(CN) <sub>2</sub>              | barium cyanide      |
| 13) Pb <sub>3</sub> N <sub>2</sub>  | lead (II) nitride         | 28) K <sub>2</sub> SO <sub>4</sub>   | potassium sulfate   |
| 14) NH <sub>4</sub> NO <sub>3</sub> | ammonium nitrate          | 29) Na <sub>3</sub> PO <sub>4</sub>  | sodium phosphate    |
| 15) FeI <sub>2</sub>                | iron (II) iodide          | 30) Fe <sub>2</sub> O <sub>3</sub>   | iron (III) oxide    |

Write the correct formula for:

- |                        |           |                                  |                   |
|------------------------|-----------|----------------------------------|-------------------|
| 1) magnesium oxide     | $MgO$     | 19) chromium(III) oxide          | $Cr_2O_3$         |
| 2) lithium bromide     | $LiBr$    | 20) gold(I) iodide               | $AuI$             |
| 3) calcium nitride     | $Ca_3N_2$ | 21) manganese(II) nitride        | $Mn_3N_2$         |
| 4) aluminum sulfide    | $Al_2S_3$ | 22) cobalt(III) phosphide        | $CoP$             |
| 5) potassium iodide    | $KI$      | 23) silver carbonate             | $Ag_2CO_3$        |
| 6) strontium chloride  | $SrCl_2$  | 24) potassium hydrogen phosphate | $K_2HPO_4$        |
| 7) sodium sulfide      | $Na_2S$   | 25) aluminum hydroxide           | $Al(OH)_3$        |
| 8) beryllium bromide   | $BeBr_2$  | 26) sodium hydrogen carbonate    | $NaHCO_3$         |
| 9) sodium oxide        | $Na_2O$   | 27) calcium acetate              | $Ca(C_2H_3O_2)_2$ |
| 10) calcium fluoride   | $CaF_2$   | 28) potassium permanganate       | $KMnO_4$          |
| 11) boron phosphide    | $BP$      | 29) calcium perchlorate          | $Ca(ClO_4)_2$     |
| 12) aluminum oxide     | $Al_2O_3$ | 30) lithium carbonate            | $Li_2CO_3$        |
| 13) iron(II) chloride  | $FeCl_2$  | 31) magnesium hydrogen sulfite   | $MgHSO_3$         |
| 14) copper(I) sulfide  | $Cu_2S$   | 32) sodium hypochlorite          | $NaClO$           |
| 15) lead(IV) iodide    | $PbI_4$   | 33) tin(IV) chlorite             | $Sn(ClO)_4$       |
| 16) tin(II) fluoride   | $SnF_2$   | 34) mercury(II) phosphate        | $Hg_3(PO_4)_2$    |
| 17) mercury(I) bromide | $HgBr$    | 35) tin(II) carbonate            | $SnCO_3$          |
| 18) tin(II) oxide      | $SnO$     |                                  |                   |

