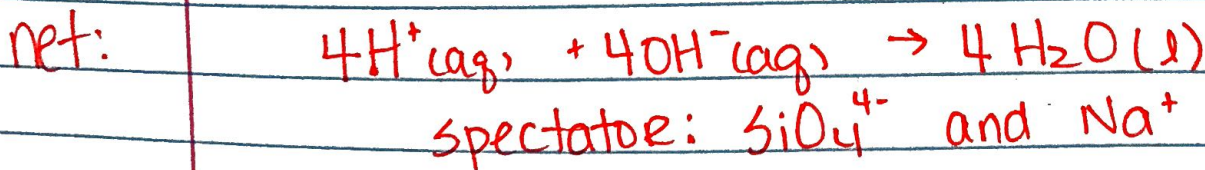
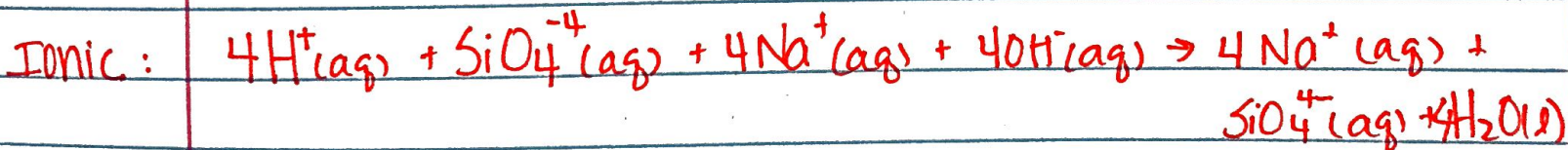
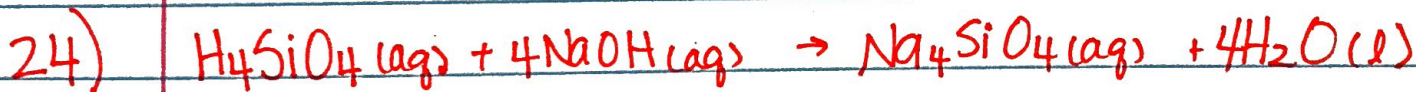
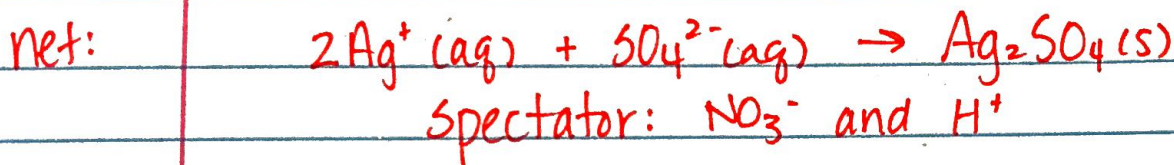
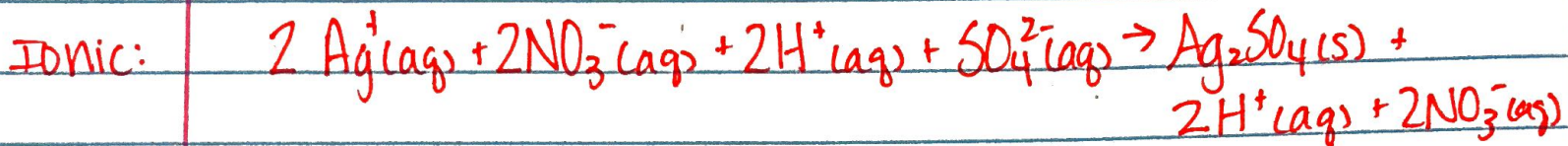
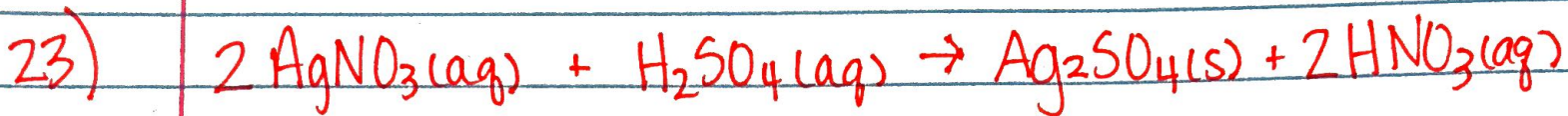


Net Ionic Equations pg. 588-589

Practice: pg 590 # 23 + 24



K_{sp} Textbook pg 602-605
Practice p. 620 # 32 + 33

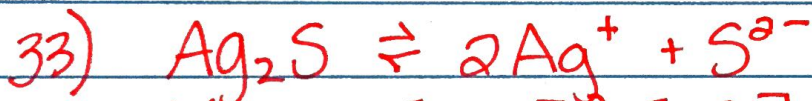


$$K_{sp} = [\text{Ca}^{2+}][\text{CO}_3^{2-}]$$

$$2.5 \times 10^{-14} = [x][x]$$

$$2.5 \times 10^{-14} = x^2$$

$$x = 1.58 \times 10^{-7} \text{ M}$$



$$K_{sp} = [\text{Ag}^+]^2 [\text{S}^{2-}]$$

← 2:1 ratio.

$[\text{Ag}^+] = 2 \times (3.4 \times 10^{-17})$ b/c there are
twice as many Ag^+ ions as S^{2-} ions
in solution

$$K_{sp} = [6.8 \times 10^{-17}]^2 [3.4 \times 10^{-17}]$$

$$K_{sp} = 1.6 \times 10^{-49}$$