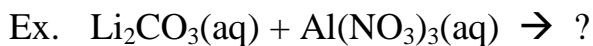
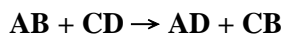


Predicting Precipitation Reactions

Directions: To predict whether a precipitation reaction will take place between two aqueous solutions and if the answer is yes, to write the complete equation for the reaction.



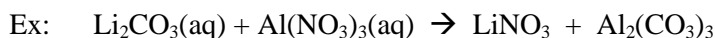
Remember: Double replacement reactions follow the general equation:



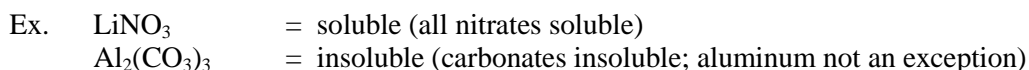
Here is an example:

STEP 1: Write the formulas for the products

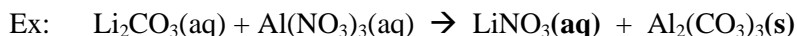
Caution: Forgetting to check the charges and criss-cross if necessary will result in an incorrect answer!



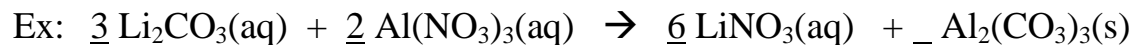
STEP 2: Use the SOLUBILITY GUIDELINES on your Reference Tables (Table F!) to determine whether either of the possible products is water **insoluble**. If either possible product is insoluble, a precipitation reaction takes place, and continue to Step 3. If neither is insoluble, write "no reaction".



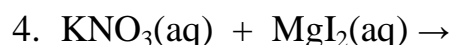
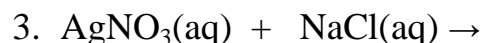
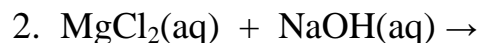
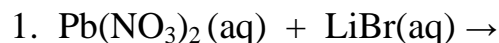
STEP 3: Put an (s) for solid next to the **insoluble** product. This is the precipitate formed in the reaction. It does not dissolve. Put (aq) next to the compound that is water-soluble.



STEP 4: Balance the equation.



Practice:



Over \rightarrow

Circle the precipitate in the following reactions.

