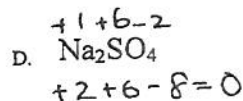
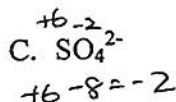
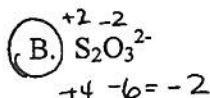
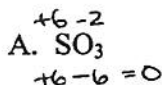


Name: _____
Redox Review Sheet

Answer Key

Directions: Select the letter of the answer choice that best completes each question. Write a statement explaining why you selected your answer choice.

1. The oxidation number of sulfur in each of the following is +6 except for:



2. Reduction is

A. A gain of electrons
 B. A loss of electrons

C. A gain of oxygen

D. Both a and c

OIL
RIG

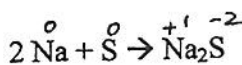
3. Identify the oxidizing agent in the following reaction.

A. Na

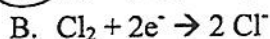
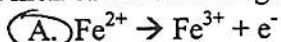
B. S

C. Na_2S

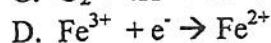
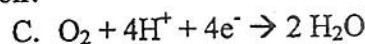
D. Na^+



4. Which of the following represents an oxidation half-reaction?



Loss of e^-



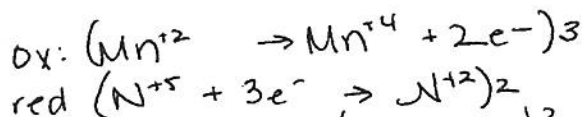
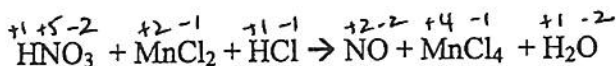
5. What will be the coefficient of HNO_3 when the following equation is completely balanced using the smallest whole number coefficients? (Hint: Use the half reaction method)

A. 2

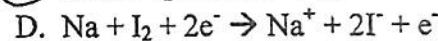
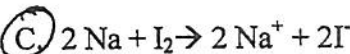
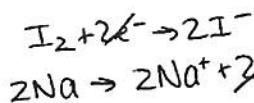
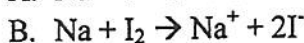
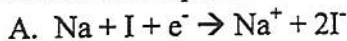
B. 3

C. 6

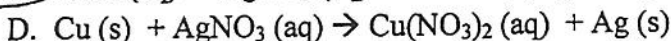
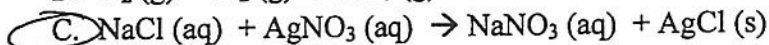
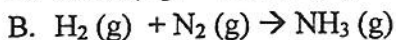
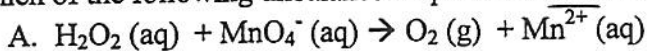
D. 5



6. When the half reactions $\text{I}_2 + 2e^- \rightarrow 2\text{I}^-$ and $(\text{Na} \rightarrow \text{Na}^+ + e^-)$ are correctly combined, the balanced redox equation is:



7. Which of the following unbalanced equation is not redox?



Double Replacement
is never redox.

8. Which of the following changes in oxidation number represents oxidation?

A. 0 to +1

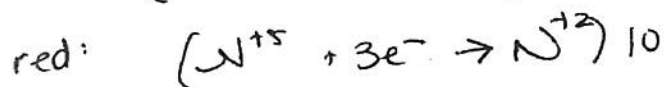
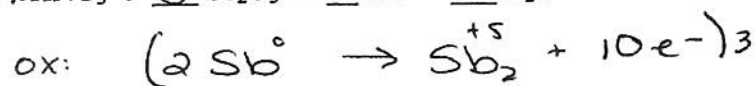
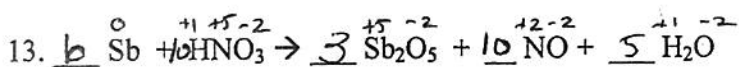
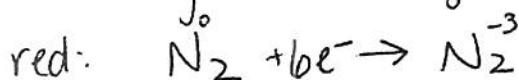
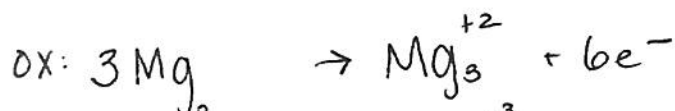
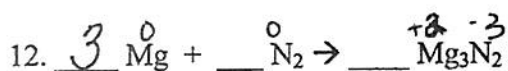
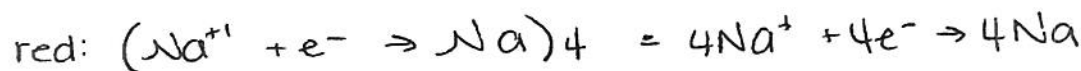
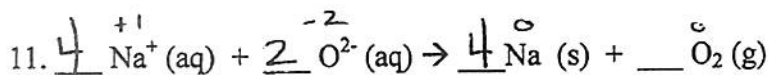
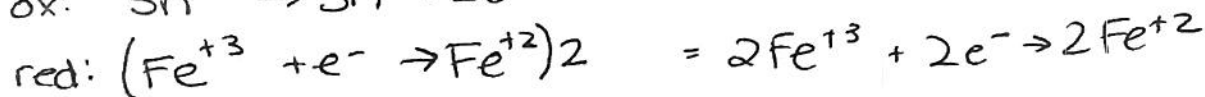
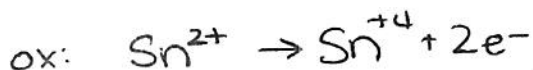
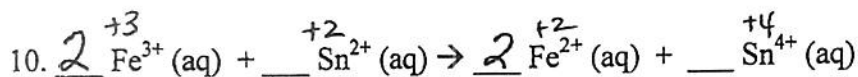
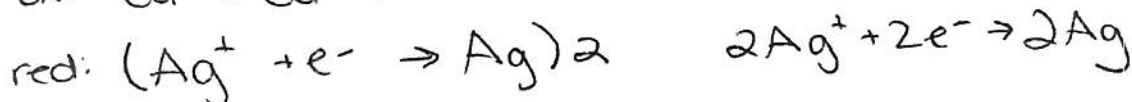
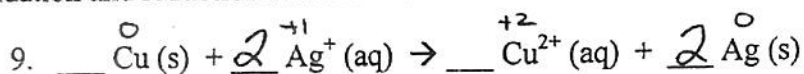
B. 0 to -1

C. +1 to -1

D. -1 to -2

— positive change
 red ↓ $\begin{array}{c} +2 \\ +1 \\ 0 \\ -1 \\ -2 \end{array}$ ↑ ox

Directions: Balance each of the following equations using the half-reaction method. Both the oxidation and reduction half-reactions must be shown.



*make sure you balance spectators.

HONORS ONLY

