

Salts, Solutions and K_{sp} Worksheet

(adapted from Glencoe Chapter 23 Study Guide and Chapter 24 Concept Review)

Name: _____

Date: _____

Key

Complete the sentence or answer the question.

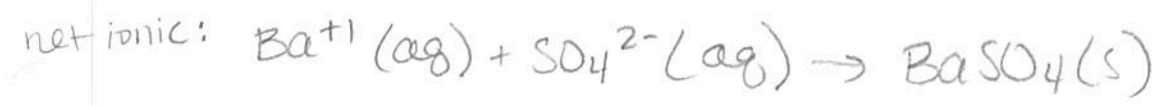
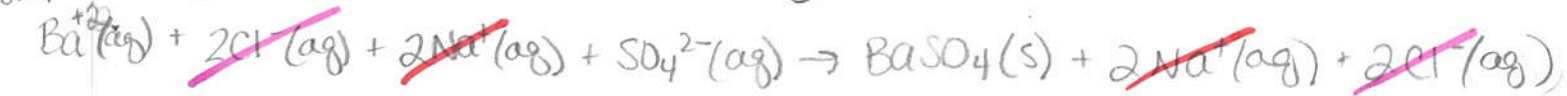
1. Ions present in a solution but not involved in the reaction are called spectator ions.

2. State the general rule that must be observed when writing net ionic equations.

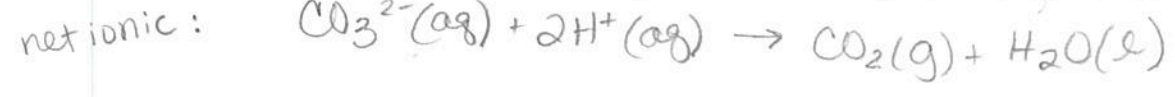
- ① Write balanced molecular equation (double replacement)
- ② Break up anything (aq) into component ions. Leave(s) intact
- ③ Eliminate spectator ions (appear same on right + left side of arrow)

3. Convert the following balanced chemical equations first into ionic form, then to net ionic form and finally identify the spectator ions.

complete ionic: a. $Ba^{+2} Cl^{-1} + Na^{+1} SO_4^{2-} \rightarrow BaSO_4(s) + 2 Na^{+1} Cl^{-1}$ leave intact ■ = cation ■ = anion



b. $Na^{+1} CO_3^{2-} + H^{+1} Cl^{-1} \rightarrow Na^{+1} Cl^{-1} + CO_2(g) + H_2O(l)$



c. $Ca(OH)_2(aq) + Na_2CO_3(aq) \rightarrow CaCO_3(s) + 2NaOH(aq)$

